

Name _____ Date _____ Period _____

Periodic Table Worksheet

Use a Periodic table to find the information asked for below:

1. What is the atomic number of:

Calcium 20
Iron 26
Gold 79
Uranium 92

2. What is the Atomic mass of:

Calcium 40.1
Iron 55.8
Uranium 238
Copper 63.5

3. How many protons do the following have?

Calcium have 20
Gold 79
Copper 29
Iron 26

4. How many electrons do the following have?

Gold have 79
Iron 26
Copper 29
Uranium 92

5. Does mercury have more protons and electrons than tin? yes

6. Is mercury a heavier element than tin? yes

7. Does potassium have more electrons than neon? yes

8. Does hydrogen have more electrons than Uranium? no

9. Which has more protons, sulfur or iodine? iodine

10. Which has more protons, iodine or silver? iodine

11. In the boxes below make Bohr models for each of the elements.

a. Determine how many electrons, protons, and neutrons there are in each atom.

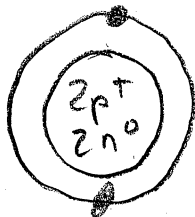
b. Draw a Bohr model of each element using the number of electrons, protons, and neutrons

c. NOTE: The first energy level can only hold up to 2 electrons. The second energy level can hold up to 8 electrons.

2 Electrons

2 Protons HELIUM

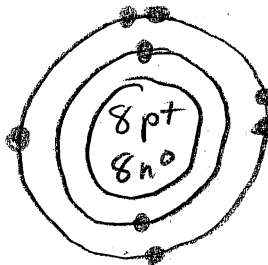
2 Neutrons



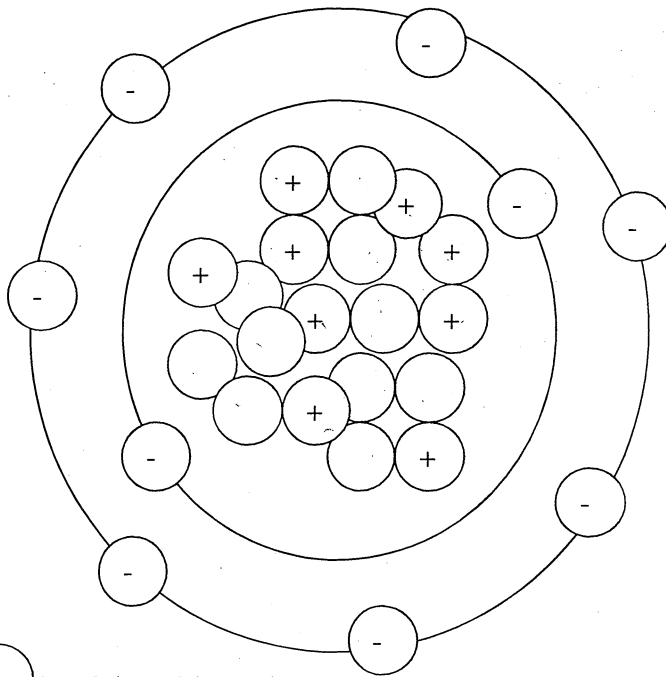
8 Electrons

8 Protons OXYGEN

8 Neutrons



12. Study the following model of an atom and answer the following questions:



Key:



Particles with no charge



Particles with negative charge



Particles with positive charge

- How many electrons does this atom have? 9
- How many protons? 9
- How many neutrons? 10
- What is the atomic number? 9
- Find the name of this element by referring to the periodic chart. Fluorine

13. Write the symbols or the names for each of these elements:

Chlorine Cl
 Copper Cu
 Potassium K
 Silver Ag
Sodium Na
Tin Sn

Zinc Zn
 Helium He
 Iron Fe
Phosphorus P
neon Ne
 Mercury Hg